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ChemActivity 6 1.The number of the valence shell corresponds to the number of the row of the Periodic Table. 2.Because the nuclei gain one proton as one moves from left to right across a period but the number of inner shell electrons remains constant.

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6.IE of Kr > IE of Br because they are in the same valence shell and Kr has the higher core charge (+8 vs. +7). IE of Rb is the lowest because core charge is +1 and its valence shell (n = 5) is larger than the valence shell (n = 4) of Kr and Br.

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6. What is the significance of the atomic number, Z, above each atomic symbol in the periodic table? 7. Based on your answer to CTQ 6, what do all nickel (Ni) atoms have in common?

ChemActivity 1 The Nuclear Atom 3

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The Nuclear Atom

Atoms, atomic number and size are factors that define an element and depending on the atomic size or number, an element can be placed as a solid, gas or liquid in a periodic table. This is a good challenge for you if you just studied the periodic table. Give it a shot!

Lesson 10 Test - Atomic Size And The Periodic Table ...

Periodic trends predict differences between elemental characteristics as you move across the periodic table. Trends are based on Coulomb's law which mathematically relates several characteristics of an elements. Atomic size measured the distance between the nucleus of an atom and the outermost non-valence electrons of the atom. Atomic size decreases from left to right, because the addition of ...

Periodic Trends in Atomic Size - Chemistry | Socratic

Answers to ChemActivity 1 - The Nuclear Atom- Chapter 5 1. 6 2. 6,7,7 3. 6,6,7 4. (a) neutral same e and p, ion different e and p (b) assign +1 to each proton and -1 to each electron and take the difference.

Chemactivity 7 Answers

Atomic Structure. 1 The Nuclear Atom 2. 2 Atomic Number and Atomic Mass 8. 3 Coulombic Potential Energy 16. 4 The Shell Model (I) 22. 5 The Shell Model (II) 30. 6 Atomic Size 40. 7 Electromagnetic Radiation 44. 8 Photoelectron Spectroscopy 48. 9 The Shell Model (III) 56. 10 Electron Configurations 60. 11 Electron Configurations and the Periodic ...

Chemistry: A Guided Inquiry, 6th Edition | Wiley

ChemActivity 6 Atomic Size 372. Why does the core charge increase as one moves from left to right across a period in the periodic table—for example, from boron to carbon to oxygen?3. What trend in atomic radius is observed as one moves from left to right across a period? Why?4.

Chemistry a Guided Inquiry Pages 1 - 50 - Text Version ...

$1\text{Mj}/10^6\text{j} \times (2.178 \times 10^{-18}\text{j})/\text{H atom} \times (6.02 \times 10^{23} \text{ H atoms})/\text{mole}$.
This = 1.31 Mj/mol Which electron has the lower ionization

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energy: the electron that is closer to the nucleus or the electron that is farther from the nucleus?

Chemactivity 4: The Shell Model (I) Flashcards | Quizlet

ChemActivity 2 Atomic Number and Atomic Mass Are All of an Element's Atoms Identical?) WARM-UP Model 1: The Average Mass of a Marble In a collection of four marbles, 25% of the marbles have a mass of 5.00 g and 75% of the marbles have a mass of 7.00 g. 5.00 g 7.00 g 26.00 g The average mass of a marble can be determined by dividing the total mass of the marbles by the total number of marbles ...

Solved: ChemActivity 2 Atomic Number And Atomic Mass Are A ...

The size of the peak in the spectrum is determined by the relative number of electrons with a given IE. Thus, the peak at 2.08 MJ/mole should be 4 times as large as the peak at a much higher energy, which corresponds to the ejection of electrons from the $n = 1$ shell. ... 66 ChemActivity 12 Electron Spin 6. a) Make a diagram for C similar to those ...

Chemistry a Guided Inquiry Pages 51 - 100 - Text Version

...

Chapter 7 - Hydrostatic Uplift Analysis 7-6 Figure 7-6 This is another example of hydrostatic pressures at an Ohio landfill causing flow through more than 20 feet of heavy in situ clay materials resulting in flooding of the excavation Note that in this case, the presence of water cannot be taken into account due to precipitation The

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